Chapter 2- TGT Questions

1. Name the two types of properties.
2. Give 2 examples of extensive properties.
3. Give 3 examples of intensive properties.
4. Why can intensive properties be used to identify a particular substance?
5. Give two examples of physical properties.
6. What are the two characteristics of a solid.
7. What are the two characteristics of a liquid.
8. What are the two characteristics of a gas.
9. Give an example of a physical change.
10. Define mixture.
11. Define phase.
12. Define substance.
13. Give 3 examples of heterogeneous mixtures.
14. Give 3 examples of homogeneous mixtures.
15. Define element.
16. Define compound.
17. Give an example of a chemical property
18. List the 4 possible clues to a chemical change (from the book, not the poster).
19. What is a precipitate?
20. What is a reactant?
21. What is a product?
22. State the Law of Conservation of Mass
23. What changes during a chemical change?
24. What do chemists use to represent elements?
25. What do chemists use to represent compounds?
26. Describe one method of separating the components of a mixture.
27. Define solution.
28. Name the two types of physical changes.
29. Give 2 examples of reversible physical changes.
30. Give 2 examples of irreversible physical changes.
31. Chemical and Physical
32. Examples could include: mass, volume, or anything else that does not depend on the amount of the substance.
33. Examples could include: density, hardness, melting temp. (freezing temp.), boiling temp. (condensing temp.), color, or anything else that doesn’t depend on the amount of the substance.
34. Every sample of a specific substance has the same intensive properties.
35. Examples include, density, hardness, melting and boiling temperatures, color, and anything that doesn’t show a substance reacting
36. Solids have a definite shape and definite volume.
37. Liquids have a definite volume, but the shape depends on the container (indefinite shape)
38. Gases do not have a definite shape or definite volume.
39. Examples include, boiling, freezing, melt, condense, break, split, grind, cut, crush or anything that does not change the nature of the substance.
40. A mixture is a physical blend of two or more components.
41. A phase is any part of a sample with uniform composition and properties.
42. A substance is a type of matter that has a uniform and definite composition.
43. Examples include: concrete, rocky-road ice cream, muddy water, or anything that has 2 or more phases
44. Examples include: salt water, homogenized milk, vanilla ice cream, mixed cool-aid, and anything that is in a single phase.
45. An element is the simplest form of matter that has a unique set of properties.
46. A compound is substance that contains two or more elements chemically combined in a fixed proportion.
47. Examples include anything that shows the substance reacting with another substance (burning, etc.)
48. The 4 clues are: a transfer of energy, a change in color, the production of a gas, the formation of a precipitate.
49. A precipitate is a solid that forms and settles out of a liquid mixture.
50. A substance present at the start of the reaction is a reactant.
51. A substance present at the end of the reaction is a product.
52. “During any chemical reaction, the mass of the products is always equal to the mass of the reactants.” Or “You cannot create or destroy mass during a chemical reaction.”
53. The compositions of the reactants always change during a chemical reaction.
54. Chemists use chemical symbols to represent elements.
55. Chemists use chemical formulas to represent compounds.
56. Examples could include, filtrating, distilling, panning, or any other method that separates mixtures on the basis of a physical property.
57. A solution is another name for a homogeneous mixture.
58. The two types of physical changes are reversible and irreversible changes.
59. Examples include state changes such as, melting, freezing, boiling and condensing, or any other change that can be ‘undone’.
60. Examples include, crushing, breaking, splitting, grinding, cutting, or any change that can’t be ‘undone’.